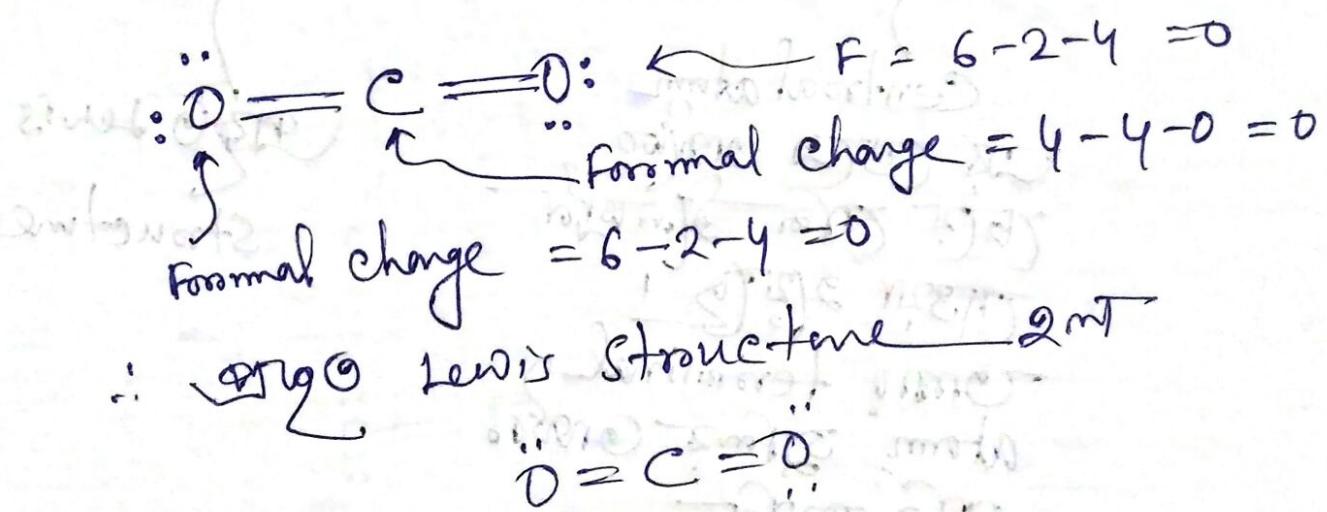




### Example - 4

$\text{ii } \text{CO}_2 \Rightarrow T = 8 \times 3 = 24$   
 $V = 4 + 6 \times 2 = 16$   
 $\therefore S = T - V = 24 - 16 = 8$   
 $\therefore \frac{S}{2} = \frac{8}{2} = 4$   
 $L = V - S = 16 - 8 = 8$   
 $\therefore \frac{L}{2} = \frac{8}{2} = 4$



### Example - 5

$$\text{CH}_2\text{N}_2 \Rightarrow T = 8 + 2 \times 2 + 8 \times 2 = 28$$
 $V = 4 + 1 \times 2 + 5 \times 2 = 16$

$$\therefore S = T - V = 28 - 16 = 12$$

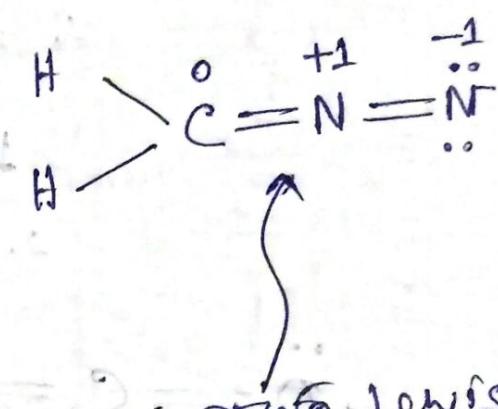
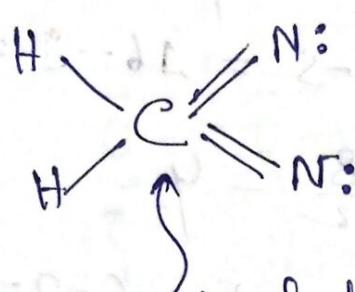
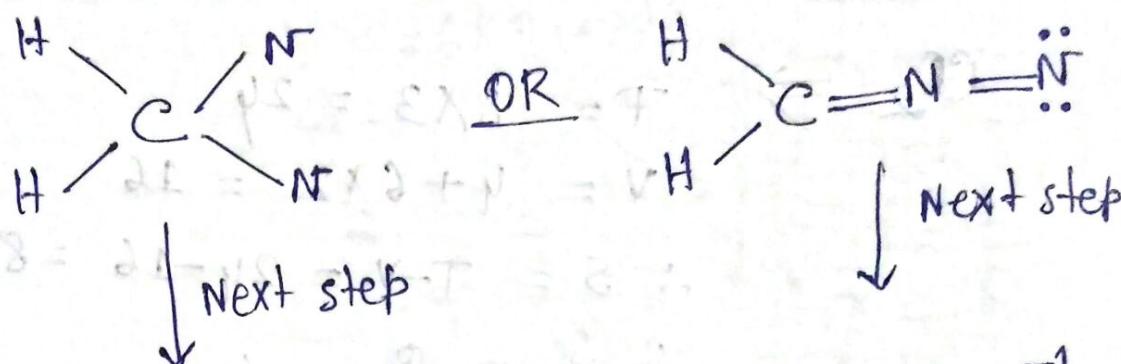
$$\therefore \frac{S}{2} = \frac{12}{2} = 6$$

$$\therefore L = V - S = 16 - 12 = 4$$

$$\therefore \frac{L}{2} = \frac{4}{2} = 2$$

(12)

Example central atom 2nd C.

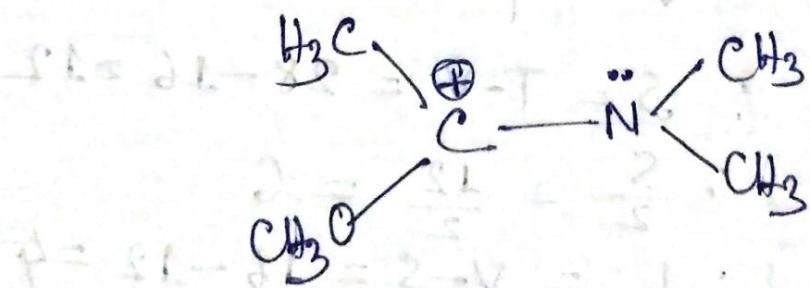
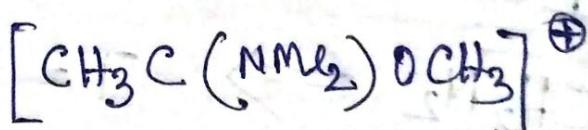


Central atom

Lewis  
Structure

Other terminal  
atom

Example - 6



### Example - 7

$$\text{NO}_3^- \Rightarrow T = 8 \times 4 = 32$$

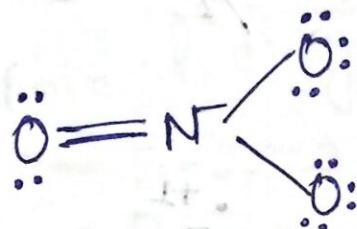
$$V = 5 + 6 \times 3 + 1 = 24$$

$$S = T - V = 32 - 24 = 8$$

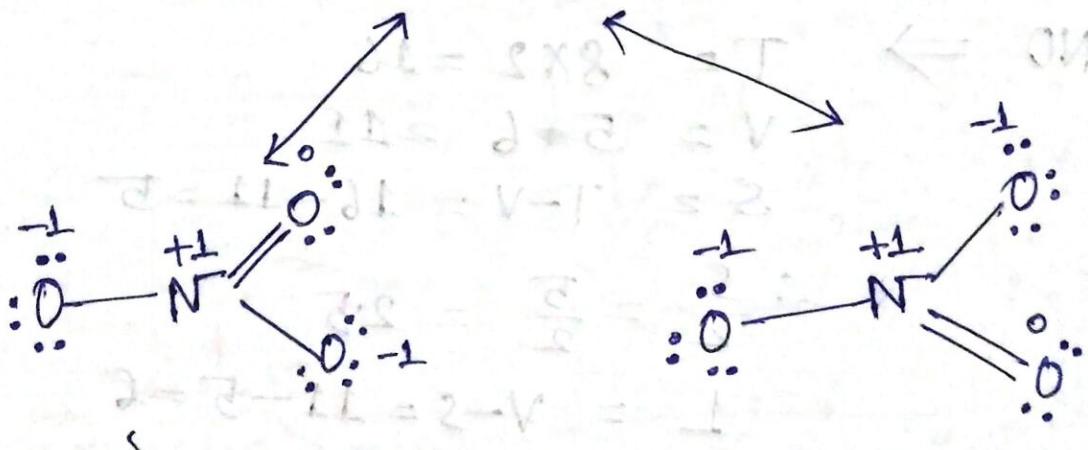
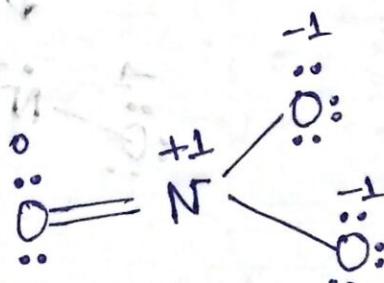
$$\therefore \frac{S}{2} = 8/2 = 4$$

$$L = V - S = 24 - 8 = 16$$

$$\therefore \frac{L}{2} = 16/2 = 8$$



Next step



Resonating structures  
 രാസത്തോട് ബന്ധപ്പെട്ട സ്ട്രക്ചർ രഹിതം

(14)

Example - 8

$$\text{NO}_2 \Rightarrow T = 8 \times 3 = 24$$

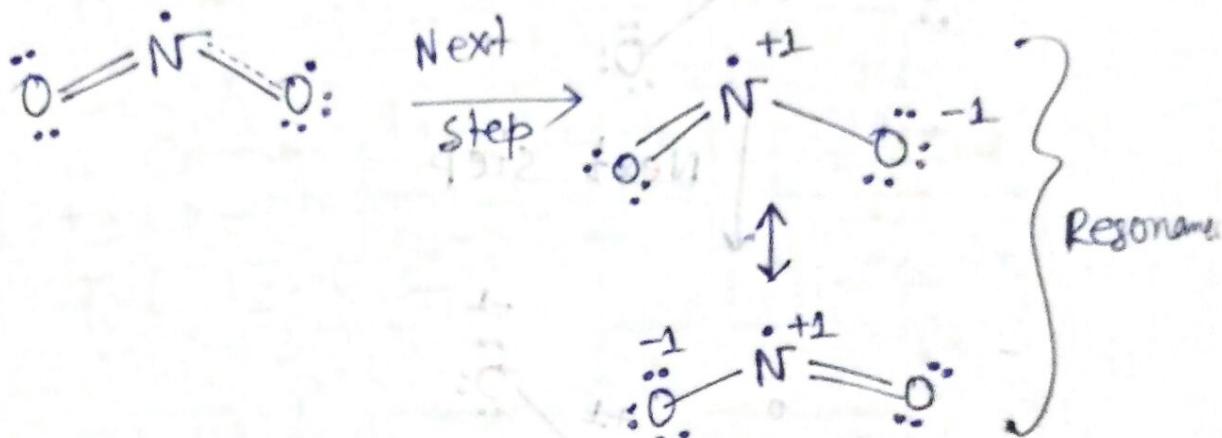
$$V = 5 + 6 \times 2 = 17$$

$$S = T - V = 24 - 17 = 7$$

$$\therefore \frac{S}{2} = \frac{7}{2} = 3.5$$

$$L = V - S = 17 - 7 = 10$$

$$\therefore \frac{L}{2} = \frac{10}{2} = 5$$

Example - 9

$$\text{NO} \Rightarrow T = 8 \times 2 = 16$$

$$V = 5 + 6 = 11$$

$$S = T - V = 16 - 11 = 5$$

$$\therefore \frac{S}{2} = \frac{5}{2} = 2.5$$

$$L = V - S = 11 - 5 = 6$$

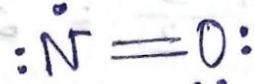
$$\therefore \frac{L}{2} = \frac{6}{2} = 3$$

Lewis Structure - 2<sup>nd</sup>

$\text{H}_2\text{O} = \text{H}_2\text{O}$



↓ Next step



Formal charge

$$= 5 - 2 - 3 \\ = 0$$

Formal charge =  $6 - 2 - 4 = 0$

Example - 10

$$\text{ClO}_4^- \Rightarrow T = 8 \times 5 = 40$$

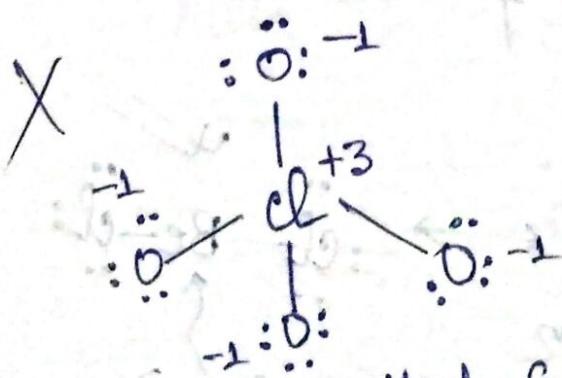
$$V = 7 + 6 \times 4 + 1 = 32$$

$$S = T - V = 40 - 32 = 8$$

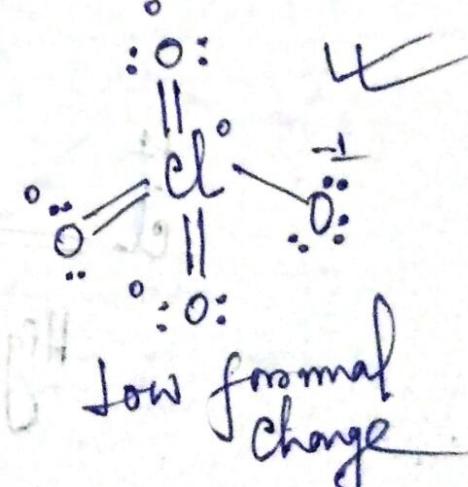
$$\therefore \frac{S}{2} = \frac{8}{2} = 4$$

$$L = V - S = 32 - 8 = 24$$

$$\therefore \frac{L}{2} = \frac{24}{2} = 12$$



High formal  
charge



Low formal  
charge

(16)

Example - 11

$$CN^- \Rightarrow T = 8 \times 2 = 16$$

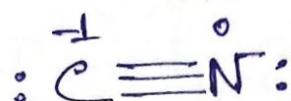
$$V = 4 + 5 + 1 = 10$$

$$S_2 = T - V = 16 - 10 = 6$$

$$\therefore \frac{S}{2} = 3$$

$$L = V - S = 10 - 6 = 4$$

$$\therefore \frac{L}{2} = 2$$

Example - 12 (Incomplete octate)

~~$BeCl_2 \Rightarrow T = 8 \times 3 = 24$~~

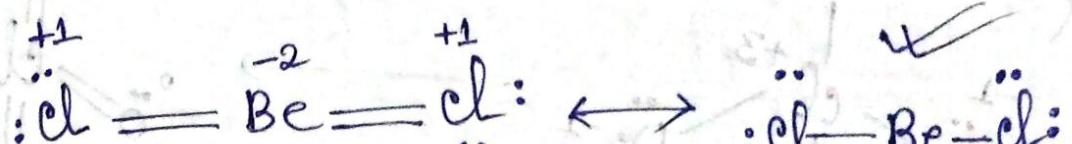
~~$V = 2 + 7 \times 2 = 16$~~

~~$S = T - V = 24 - 16 = 8$~~

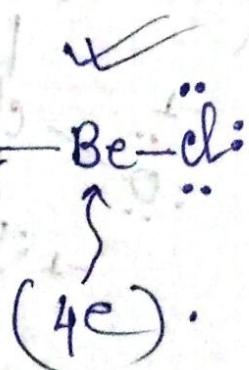
~~$\therefore \frac{S}{2} = \frac{8}{2} = 4$~~

~~$L = V - S = 16 - 8 = 8$~~

~~$\therefore \frac{L}{2} = 4$~~



~~High formal charge  
(Wrong str.)~~



Example - 13

$$BF_3 \Rightarrow T = 8 \times 4 = 32$$

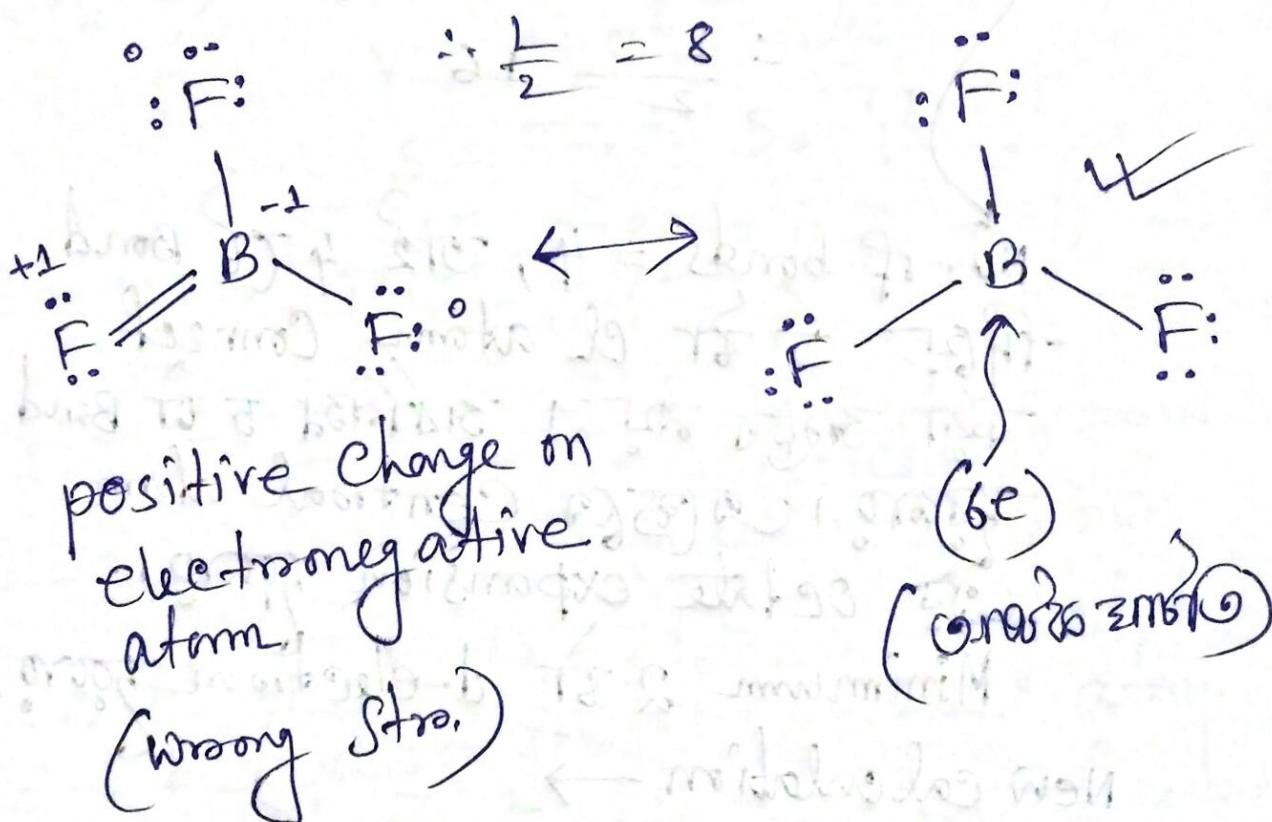
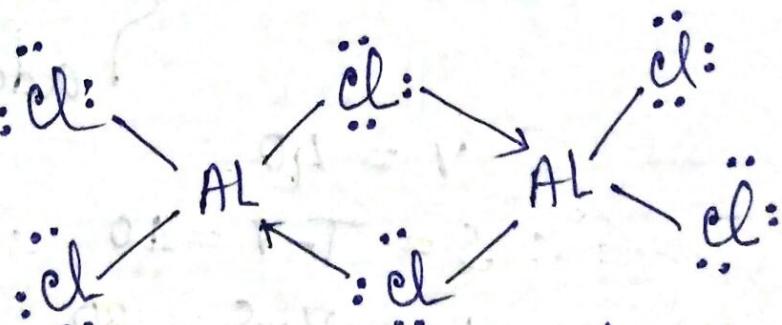
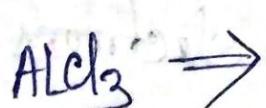
$$V = 3 + 7 \times 3 = 24$$

$$S = T - V = 32 - 24 = 8$$

$$\therefore \frac{S}{2} = 4$$

$$\therefore L = V - S = 24 - 8 = 16$$

~~Am (2235)~~

Example - 14

Dimers formation to  
fulfill octate of each Al

(18)

## Example of octate Expansion

(জনতাপত্র)

$$\textcircled{1} \quad \text{PCl}_5 \Rightarrow \text{EXT} = 8 \times 6 = 48$$

$$V = 5 + 7 \times 5 = 40$$

$$S = T - V = 48 - 40 = 8$$

$$\therefore \frac{S}{2} = 4$$

$$L = V - S = 40 - 8 = 32$$

$$\therefore \frac{L}{2} = 16$$

No. of bonds = 4, এর 4 CB Bond

প্রথমে 5 টি Cl atom connect  
যাতে 3 মধ্যে নতুন 5 টি Bond

দূরবর্তী 1 Central atom

অক্ষেভাবে Octate expansion হওয়া।

Minimum 2 টি d-electron হওয়া।

New calculation  $\rightarrow$

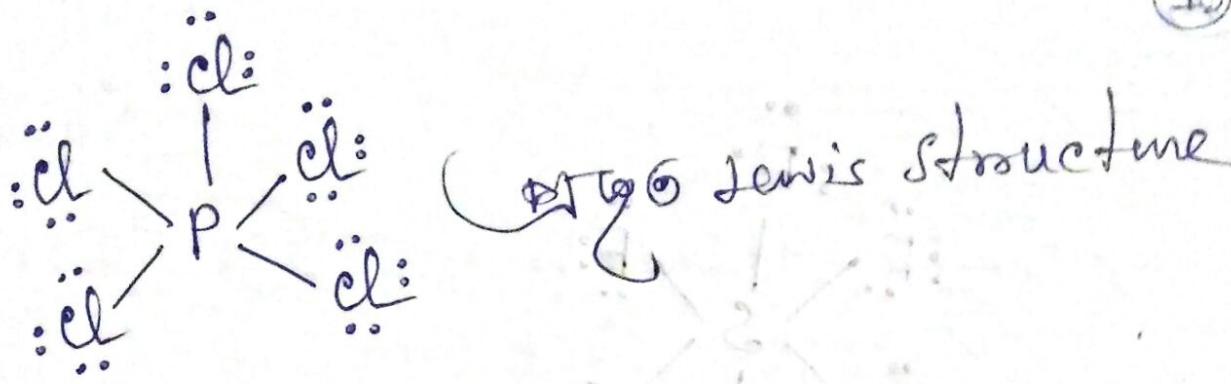
$$T = 8 \times 6 + 2 = 50$$

added electrons

$$V = 40$$

$$\therefore S = T - V = 10 \quad \therefore S/2 = 5$$

$$L = V - S = 30 \quad \therefore L/2 = 15$$



### Example - 2

$$\text{SF}_6 \Rightarrow T = 7 \times 8 = 56$$

$$V = 6 + 7 \times 6 = 48$$

$$S = T - V = 56 - 48 = 8$$

$$\therefore \frac{S}{2} = \frac{8}{2} = 4$$

insufficient to form  
at least six bonds.

- ∴ At least, 4 d-electrons need to be added for the central atom.
- ∴ octate expansion of central atom.

$$\text{New calculation} \rightarrow T = 56 + 4 = 60$$

$$V = 48$$

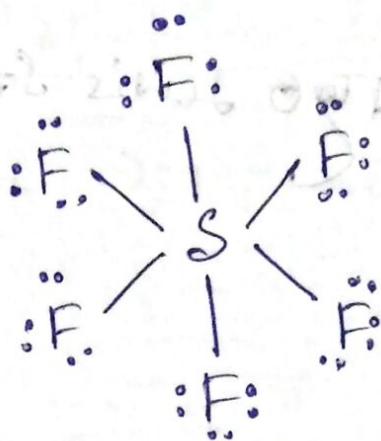
$$S = 60 - 48 = 12$$

$$\therefore \frac{S}{2} = 6$$

$$L = V - S = 48 - 12 = 36$$

$$\therefore \frac{L}{2} = 18$$

(20)

Example - 3

$$\text{IO}_2\text{F}_2^{\ominus} \Rightarrow S = T - V = 40 - 34 = 6$$

$$\therefore \frac{S}{2} = 3$$

$\therefore$  2 electrons need to be added.

$$\text{New calculation} \rightarrow T = 40 + 2 = 42$$

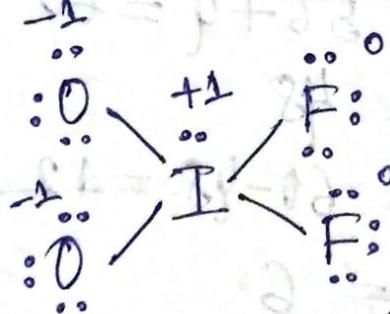
$$V = 34$$

$$S = 42 - 34 = 8$$

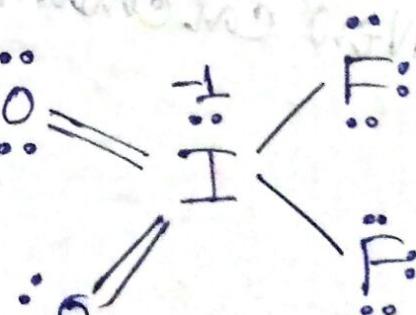
$$\frac{S}{2} = 4$$

$$L = V - S = 34 - 8 = 26$$

$$\therefore \frac{L}{2} = 13$$



High formal charge  
(wrong str.)



(Final str.)